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**PRECIPITATION IN POWDER-METALLURGY,
NICKEL-BASE SUPERALLOYS: REVIEW OF
MODELING APPROACH AND FORMULATION OF
ENGINEERING (POSTPRINT)**

**S.L. Semiatin
AFRL/RX**

**F. Zhang
CompuTherm**

**R. Larsen
TPRL**

**L.A. Chapman
NPL**

**D.U. Furrer
Pratt & Whitney**

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MATERIALS AND MANUFACTURING DIRECTORATE
WRIGHT-PATTERSON AIR FORCE BASE, OH 45433-7750
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6. AUTHOR(S)

- 3) R. Larsen - TPRL
- 4) L.A. Chapman - NPL
- 5) D.U. Furrer - Pratt & Whitney

7. PERFORMING ORGANIZATION NAME(S) AND ADDRESS(ES)

- 3) Thermophysical Properties Research Laboratory, Inc.,
3080 Kent Ave, West Lafayette, IN 47906
- 4) National Physical Laboratory, Teddington,
Middlesex TW11 0LW, UK
- 5) Pratt & Whitney, 400 Main Street,
East Hartford, CT 06118

RESEARCH

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Precipitation in powder-metallurgy, nickel-base superalloys: review of modeling approach and formulation of engineering methods to determine input data

S. L. Semiatin^{1*}, F. Zhang², R. Larsen³, L. A. Chapman⁴ and D. U. Furrer⁵

* Correspondence:

sheldon.semiatin@us.af.mil

¹Air Force Research Laboratory,
Materials and Manufacturing
Directorate, AFRL/RXCM,
Wright-Patterson Air Force Base, OH
45433-7817, USA

Full list of author information is
available at the end of the article

Abstract

Methods for determining the various thermodynamic and kinetic parameters required for the modeling of γ' precipitation in powder-metallurgy (PM), nickel-base superalloys are summarized. These parameters comprise the composition of the γ' phase, the γ' solvus temperature/equilibrium solvus approach curve, the free energy (ΔG^*) associated with the decomposition of the γ matrix to form γ' , the γ/γ' interfacial energy σ , and an effective diffusivity for use in nucleation, growth, and coarsening calculations. Techniques to obtain the material data include phase extraction (for the average composition of γ') and heat-treatment/quantitative metallography (for a two-parameter fit of the solvus approach curve). With regard to ΔG^* , two methods, one based on the instantaneous composition of the γ and γ' phases and the other on the enthalpy of transformation and the solvus temperature, are summarized. It is shown that the interfacial energy σ can be determined from the nucleation-onset temperature as indicated by on-cooling specific-heat measurements. Last, the use of a limited set of static-coarsening measurements to estimate the effective diffusivity is described. The application of the various protocols is illustrated for typical first-, second-, and third-generation PM superalloys, i.e., IN-100, René 88, and LSHR/ME3, respectively.

Keywords: Precipitation, Superalloys, Gamma prime, Nucleation, Growth, Coarsening, Solvus approach curve, Interface energy, Free energy of transformation, Diffusivity

Background

Nickel-base superalloys comprise an important class of high-temperature structural materials used widely in jet-engine and land-based, power-generation applications [1]. Depending on alloy content and intended service, superalloys are typically synthesized via ingot metallurgy (IM; aka cast + wrought), powder-metallurgy (PM), or investment-casting (IC) methods [1, 2]. Lean alloys, which are not prone to macro-segregation or thermal cracking in ingot form, are usually made via the first technique. PM methods are employed to manufacture semi-finished billet products and are common for more-highly alloyed materials prone to segregation or thermal cracking when cast in large diameters. Investment casting is frequently applied for the net-shape manufacture of components of the most-highly alloyed compositions. Irrespective of processing method, however, all

superalloys rely on various combinations of precipitation and solid-solution strengthening of the face-centered cubic (fcc) gamma (γ) matrix to maintain high-temperature strength.

The principal strengthening precipitates in nickel-base superalloys consist of the ordered fcc gamma-prime (γ') and ordered body-centered tetragonal (bct) gamma-double-prime (γ'') phases [1–3]. The former are commonly found in all three types of alloys (IM, PM, IC) and have a composition of Ni_3X , in which $\text{X} = \text{Al}, \text{Ti}, \text{Ta}, \text{and Nb}$. The latter precipitates are most common in IM materials such as superalloy 718 and have a composition of Ni_3Nb . Alloy 718 often has additional, coarser Ni_3Nb delta (δ) precipitates (with an *orthorhombic* crystal structure) that form at higher temperatures and are used primarily to control the size of the γ grains.

Processing temperatures for nickel-base superalloys are usually chosen relative to the precipitate solvus (or solvi, as is the case for 718), i.e., the temperature(s) above which all precipitates are dissolved. For example, following powder synthesis and consolidation via blind-die compaction or hot isostatic pressing, γ' -strengthened PM superalloys are typically extruded at a subsolvus temperature to develop a fine, recrystallized billet microstructure composed largely of γ grains and *primary* γ' precipitates. The average diameter in each phase is of the order of 1 to 5 μm . Under low-strain-rate processing conditions in the two-phase field, alloys with this microstructure are superplastic. Sections cut from the billet product are thus readily shaped via isothermal-forging techniques.

Depending on the service application, subsequent heat-treatment conditions are also selected relative to the solvus. For applications requiring high strength and resistance to fatigue-crack initiation, heat treatment for γ' -strengthened alloys consists of solution treatment below the γ' solvus, free- or forced-air cooling to room temperature, and final aging. The γ' precipitates formed from such a heat treatment are denoted as *primary* (retained from prior subtransus processing during soaking at the solution temperature), *secondary* (formed at higher temperatures during cooling following solution treatment), and *tertiary* (formed at lower temperatures during cooldown and during final aging). For applications requiring good creep strength and crack-growth resistance, a solution heat treatment consisting of supersolvus exposure to grow the γ grain size is typical. In this case, secondary and tertiary γ' precipitates are formed during cooling.

Because of the importance of precipitation to the mechanical properties, it is not surprising that considerable work has been performed to model it in terms of pertinent nucleation, growth, and coarsening phenomena [4–24]. A number of efforts have focused on γ - γ' superalloys in particular, primarily for the isothermal heat treatment of alloys with 2, 3, or 4 components [25–29]. Several investigations have also examined the more complex case of continuous cooling of multi-component commercial alloys [30–36]. Furthermore, classical deterministic analyses of precipitation have been complemented by more-detailed statistical treatments such as those based on the phase-field method, e.g., references [37–41]. Although computationally more intensive, these approaches are often used to quantify those conditions under which precipitates evolve from a spheroidal to a cuboidal shape as they grow large. Thus, they can provide better quantitative predictions of the size of large precipitates than classical analyses that assume spheroidal growth only.

The objective of the present work was to develop engineering methods for determining the input material data for classical models of *intragranular* γ' precipitation,¹ especially those parameter descriptions that can be utilized readily in fast-acting numerical

simulations. The efficacy of the approaches was demonstrated for several commercial PM superalloys. To provide context for this work, the various models used to describe precipitation are summarized first.

Fundamentals of precipitation modeling

Typical modeling procedures for precipitation are based on classical relations for the rate of homogeneous nucleation, diffusional growth, and static coarsening.

Nucleation rate

The rate of nucleation of precipitates, J , is given by the following equation [4, 12, 26]:

$$J = \frac{2CD}{a_o^4} \sqrt{\frac{\sigma}{k_B T}} \cdot \exp\left(-\frac{4\pi\sigma r^{*2}}{3k_B T}\right) \cdot \exp(-\tau/t) = J_o \exp(-\tau/t), \quad (1)$$

in which C denotes the solute content in the matrix expressed as an atomic fraction, D is the solute diffusivity in the matrix phase, a_o is the average lattice parameter of the matrix and precipitate phases (taken to be ~ 0.356 nm for most γ - γ' superalloys), σ is the matrix-precipitate surface energy, k_B is Boltzmann's constant (1.3806 JK^{-1}), T is the temperature in Kelvin, and t is the time. The critical radius of the precipitate, r^* , is a function of σ , the chemical free energy of transformation, ΔG^* (taken to be a positive quantity for a decrease in energy), and the elastic strain energy, ΔG_p , associated with the difference in lattice parameters of the matrix and precipitate phases, i.e.,

$$r^* = \frac{2\sigma}{\Delta G^* + \Delta G_p}. \quad (2)$$

The nucleation rate J in Eq. (1) is related to the steady-state nucleation rate J_o through the term $\exp(-\tau/t)$, which describes the initial nucleation transient during which a meta-stable distribution of embryos with sub-critical radii are formed. The so-called incubation time constant τ in this exponential term is given by the relation [26, 42]:

$$\tau = \frac{\pi R T r^{*3}}{96 V_M D \sigma}, \quad (3)$$

in which R is the gas constant, and V_M is the molar volume of the precipitate.

Strictly speaking, Eqs. (1)–(3) apply to nucleation in *two-component* alloys. Methods and assumptions utilized to determine *effective* values for coefficients such as C , D , and ΔG^* of multi-component alloys for use in the nucleation (as well as growth and coarsening) relations are discussed below.

Diffusional growth

In most theoretical treatments for γ' precipitates, the particles are assumed to be spherical of radius r with their growth controlled by diffusion. The “exact” solution of the diffusion equation for the rate of growth in the presence of a finite matrix supersaturation is as follows [13, 14]:

$$dr/dt = 2\lambda^2 D/r, \quad (4)$$

in which λ^2 is related to the supersaturation Ω by the following relation:

$$\{\lambda^2 \exp(\lambda^2)\} \cdot [(\exp(-\lambda^2)) - (\lambda \pi^{1/2} \operatorname{erfc}(\lambda))] = \Omega/2. \quad (5)$$

The supersaturation Ω has its usual definition, i.e.,

$$\Omega = (C_m - C_i)/(C_p - C_i). \quad (6)$$

Here, C_m , C_i , and C_p represent the compositions of the matrix far from the matrix-particle interface, the matrix at the matrix-particle interface, and the particle at the matrix-particle interface, respectively. The correction to the equilibrium ($r = \infty$) interface composition, C_i , due to the Gibbs-Thompson effect can be quantified using the following equation [43]:

$$C_i(r) = C_i(r = \infty) \left(1 + \frac{2(1 - C_m)\sigma V_M}{(C_p - C_m)RT r} \right). \quad (7)$$

The effect of soft impingement on the “far-field” matrix composition C_m is taken into account using the usual approximation derived from a mass balance [44]:

$$C_m = (C_o - fC_p)/(1 - f), \quad (8)$$

in which C_o and f denote the overall alloy composition and the volume fraction of the precipitate.

As for the nucleation-rate relation (Eq. (1)), the diffusional-growth solution (Eq. (4)) is strictly applicable to a two-component system whose inter-diffusion coefficient is independent of composition. For multi-component alloys, D thus represents again an effective diffusivity (D_{eff}). Furthermore, it is often assumed in fast-acting simulations that the composition of the γ' precipitate is constant and equal to the average value determined by phase extraction.

Coarsening

During continuous cooling at rates typical of production-scale components, the tendency for coarsening of γ' precipitates (especially the larger secondary precipitates) is small due to retained supersaturation. On the other hand, coarsening during aging (and during service at high temperatures) can be quite substantial. In the present work, measurements of the isothermal, static-coarsening rate have been used to estimate the effective diffusivity needed in the expressions for both nucleation and growth. Specifically, the rate of increase of the average precipitate size was interpreted in terms of the classical Lifshitz-Slyozov-Wagner (LSW) theory modified to treat the coarsening of finite volume fractions of particles (15, 16, 21), viz.:

$$\bar{r}^3 - \bar{r}_o^3 = K_{\text{MLSW}} t \quad (9)$$

and

$$K_{\text{MLSW}} = \frac{8w(f)D\sigma C_\gamma(1 - C_\gamma)V_M}{9RT(C_{\gamma'} - C_\gamma)^2 [1 + \partial \ln v / \partial \ln C_\gamma]}. \quad (10)$$

In Eq. (9), \bar{r} and \bar{r}_o represent the *average* instantaneous and initial particle radii. In the expression for the modified LSW rate constant, K_{MLSW} , terms not previously defined include $w(f)$, the factor to correct for the finite volume fraction of particles [17–20], and C_γ and $C_{\gamma'}$, the equilibrium concentrations of the rate-limiting solute in the matrix and

precipitate, respectively. The bracketed term in the denominator of Eq. (10) is the thermodynamic factor in which ν denotes the activity coefficient for the rate-limiting solute in the γ matrix of specified composition.

Again, the formulation and application of Eqs. (9) and (10) for multi-component alloys implicitly assumes that coarsening is controlled by the diffusion of a single, rate-limiting solute. However, Kuehmann and Voorhees [24] demonstrated that rapidly diffusing solutes can also affect the coarsening rate and the related diffusional processes for *ternary* alloys. For multi-component alloys, their work suggests that an effective rate constant (K_{eff}) equal to the inverse of the sum of the inverse rate constants for the individual solutes can be defined as follows:

$$\frac{1}{K_{\text{eff}}} = \frac{1}{K_1} + \frac{1}{K_2} + \frac{1}{K_3} + \dots \quad (11)$$

Model input data: approaches and examples

The key material parameters for modeling of precipitation belong to one of three sets: (i) thermodynamic data, (ii) driving/retarding forces for nucleation, and (iii) kinetic factors/diffusivity. Each of these is discussed in more detail below with pertinent results for the first-, second-, and third-generation superalloys IN-100, René 88, and LSHR/ME3, respectively.

Thermodynamic data

The key thermodynamic data consist of (1) the solvus approach curve (equilibrium fraction of γ' as a function of temperature, $f(T)$, for γ - γ' superalloys), (2) the composition of the γ' phase, and (3) the solute content in the matrix (C in Eq. (1)).

Two methods exist for determining the solvus approach curve. The first is the formal Calphad (thermodynamic) method based on the calculation of the fractions of γ and γ' as a function of temperature using the Gibbs free energies of these phases [45]. For multi-component alloys such as commercial nickel-base superalloys, these calculations require an appropriate thermodynamic database typically obtained through the optimization of a number of model parameters from various measured binary, ternary, etc. phase diagrams. An alternate, semi-empirical method was developed by Dyson [46] and refined by Payton [47]. Specifically, it comprises fitting $f(T)$ measurements to an analytical expression, whose application is especially attractive for fast-acting simulations of precipitation kinetics, viz.,

$$f = - \frac{4C^* \left(1 - \exp \left[\frac{Q}{R} \left(\frac{T - T_{\gamma'}}{T_x T_{\gamma'}} \right) \right] \right)}{1 - 4C^* \exp \left[\frac{Q}{R} \left(\frac{T - T_{\gamma'}}{T_x T_{\gamma'}} \right) \right]}. \quad (12)$$

Here, $T_{\gamma'}$ denotes the γ' solvus temperature, C^* is the atomic fraction of γ' formers in the alloy, and Q is a fitting parameter. Because the γ' phase has a composition of Ni_3X , $4C^*$ is approximately equal to the maximum volume fraction of γ' in the alloy, i.e., the amount at 0 K. The value of the fitting parameter Q typically lies in the range of 60 to 75 kJ/mol.

The applicability of Eq. (12) is illustrated using IN-100, René 88, and LSHR as model alloys. From heat treatments at various temperatures followed by metallography, the solvus temperature for the three materials has been determined to be 1458, 1378,

1430 K, respectively. Typical compositions of the alloys are summarized in Table 1 (in weight percent) and Table 2 (in atomic percent). For the superalloys of interest here, the principal γ' -forming elements are Al, Ti, Nb, and Ta. From Table 2, the values of $4C^*$ are thus 0.60, 0.38, and 0.53. These estimates are in reasonable agreement with published reports of the fraction of γ' in fully heat-treated materials except for René 88, for which the measured volume fraction is 0.42, rather than 0.38 [30]. Differences between estimated and published magnitudes of the volume fraction of precipitate can be largely ascribed to the fact that the various γ' -forming elements do not partition totally to this phase, and the other alloying elements do not fully partition to the γ matrix. The specific alloying-element partitioning is illustrated by measurements of the average composition of γ' determined by phase extraction (Table 3) [48]. This method weights the composition in favor of those precipitates with the larger volume fractions, i.e., the primary and secondary γ' , of course.

Using the above values of $T_{\gamma'}$ and $4C^*$, measurements of the equilibrium volume fractions of γ' at various temperatures for each alloy have been fitted to Eq. (12) (Fig. 1). The measurements were performed via quantitative analysis of SEM micrographs taken on samples that had been heat treated for long times (4–24 h, depending on temperature) or isothermally forged to achieve equilibrium conditions. The best fits are obtained for $Q = 75$, 65, or 60 kJ/mol for IN-100, René 88, and LSHR, respectively. In Fig. 1b, the René 88 measurements and the fit using Eq. (12) also show good agreement with the solvus approach curve determined by a Calphad-type approach [49]. In addition to the precise shape of the free-energy curves as a function of temperature, the variation in the value of Q for the three alloys may be due to uncertainties in the measured volume fractions of γ' used for fitting (approximately ± 0.015) and the precise values of the maximum volume fraction of γ' (assumed to be $4C^*$).

The solute content in the matrix (C in the nucleation Eq. (1)) depends on the volume fraction of precipitate and its composition. For supersolvus solution treatments, the majority of the secondary- γ' nuclei form during a short time interval during which the matrix composition undergoes little change. Thus, C is relatively constant and equal to ~ 0.45 – 0.50 (Table 2). Furthermore, a sensitivity analysis in reference [50] has shown that C equal to 0.25, 0.5, or even the physically unrealistic value of 1.0 all yield similar precipitation predictions.

Driving force for nucleation

The principal driving force for nucleation is the reduction in chemical free energy, ΔG^* , associated with the formation of γ' from a supersaturated γ matrix. In addition to the formal Calphad approach, two somewhat simpler engineering methods can be used to determine ΔG^* [50]. One is based on measurements of the specific heat and the γ'

Table 1 Alloy compositions (weight percent)

Material	Co	Cr	Al	Ti	Nb	Ta	Mo	W	Zr	V	Ni
IN-100	18.5	12.4	5.0	4.3	–	–	3.2	–	0.06	0.8	Bal.
René 88	13	16	2.1	3.7	0.7	–	4	4	0.05	–	Bal.
LSHR	20.4	12.3	3.54	3.45	1.49	1.52	2.71	4.28	0.05	–	Bal.
ME3	20.7	12.7	3.4	3.7	0.9	2.5	3.8	2.1	0.05	–	Bal.

Table 2 Alloy compositions (atomic percent)

Material	Co	Cr	Al	Ti	Nb	Ta	Mo	W	Zr	V	Ni
IN-100	17.2	13.1	10.1	4.92	–	–	1.83	–	0.04	0.86	Bal.
René 88	12.9	17.9	4.5	4.5	0.44	–	2.4	1.3	0.03	–	Bal.
LSHR	20.1	13.7	7.61	4.18	0.93	0.49	1.64	1.35	0.03	–	Bal.
ME3	20.2	14.1	7.27	4.46	0.56	0.80	2.28	0.66	0.03	–	Bal.

solvus temperature. The other involves a thermodynamic calculation based on a pseudo-binary of nickel and chromium.

The first technique for determining ΔG^* is strictly applicable to the matrix composition pertaining to supersolvus solution treatment, i.e., the overall alloy composition. Assuming that the enthalpy (ΔH_{avg}) and entropy (ΔS_{avg}) of formation (per mol) are constant, ΔG^* varies linearly with temperature, i.e.,

$$\Delta G^*(\gamma \rightarrow \gamma') = \Delta H_{\text{avg}} - T \Delta S_{\text{avg}}. \quad (13)$$

The enthalpy of formation of the precipitate phase, ΔH_{avg} , is determined from the integral of the measured specific heat (C_p) over the *on-cooling*-transformation-temperature range (typically ~30 K below the solvus \rightarrow ~1150 K for PM nickel-base superalloys). This integral is *decremented* by the heat content associated with temperature changes in the absence of transformation as quantified by extrapolations from low and high-temperature C_p behaviors, respectively. A schematic illustration of the construction, in which the enthalpy of transformation is shown as the cross-hatched region, is shown in Fig. 2. Knowledge of the fraction transformed (f_{trans}) yields the desired quantity:

$$\Delta H_{\text{avg}} = \left[\left(\int C_p dT \right) - \text{Baseline} \right] / f_{\text{trans}} \quad (14)$$

At the γ' solvus, ΔG^* is equal to zero. Thus, the value of ΔS_{avg} is equal to $\Delta H_{\text{avg}}/T_{\gamma'}$.

Specific-heat data measured during cooling from above the solvus at a rate of 20 K/min reveal a similar dependence on temperature for IN-100, René 88, and LSHR (Fig. 3) [51–54]. Results for ME3, an alloy whose overall composition and γ' composition are similar to those of LSHR (Tables 1, 2, and 3), are also included in Fig. 3c. All of the measurements indicate a finite undercooling below the solvus at which the transformation begins, a peak just below this onset temperature, and a rapid decrease in the magnitude of C_p thereafter. The rate of decrease after the peak can be ascribed to the decreasing rate of formation of γ' inasmuch as the instantaneous fraction of precipitate follows the solvus approach curve, at least approximately, once nucleation starts to occur [50]. The estimates of ΔH_{avg} and ΔS_{avg} (Table 4) show very similar results for IN-100 and LSHR, but a measurable difference for René 88.

Table 3 Gamma-prime compositions (atomic percent) determined by phase extraction [48]

Material	Co	Cr	Al	Ti	Nb	Ta	Mo	W	Zr	Ni
IN-100	10.3	3.5	14.1	8.04	–	–	0.89	–	0.02	Bal.
René 88	4.36	4.18	9.74	11.70	1.11	–	1.55	1.68	0.06	Bal.
LSHR	11.2	2.12	12.05	7.86	1.61	0.96	0.75	1.24	0.06	Bal.
ME3	10.1	2.02	12.06	8.33	0.98	1.42	0.83	0.53	0.05	Bal.

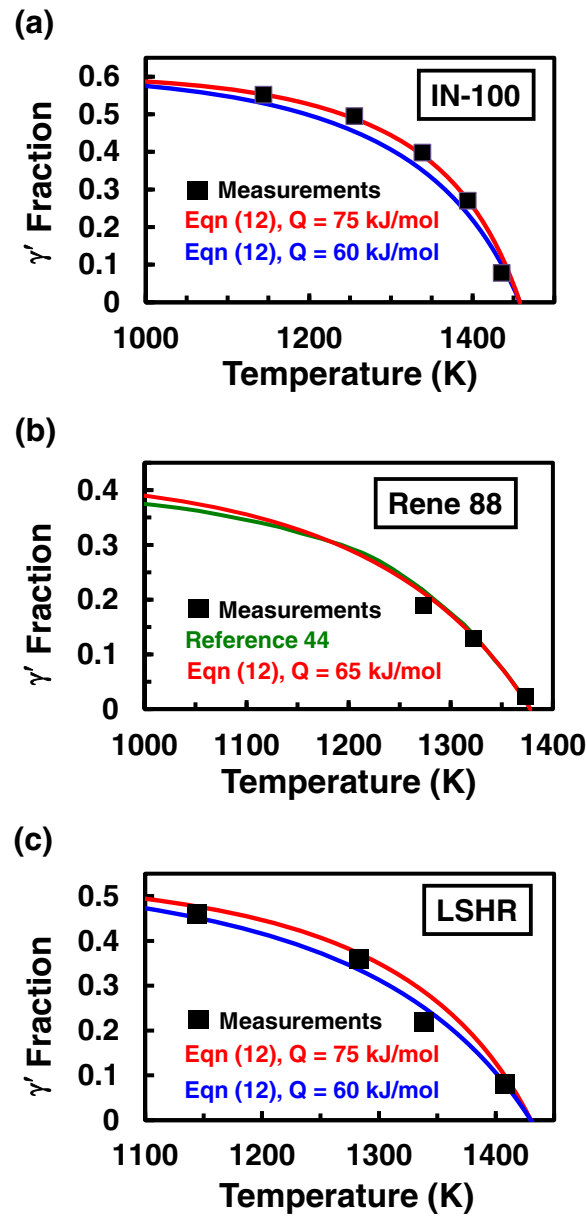


Fig. 1 Comparison of measured values of the volume fraction of γ' and analytical fits based on Eq. (12) for **a** IN-100, **b** René 88, and **c** LSHR. The results in **b** are also compared to a Calphad (thermodynamic) calculation from reference [49]

The second approach used to quantify ΔG^* is based on classical expressions from solution thermodynamics. For the case in which the precipitate is *enriched* in solute (which is applicable for Al, Ti, Nb, Ta), ΔG^* is given by [4]:

$$\Delta G^*(\gamma \rightarrow \gamma') = -\frac{(C_{\gamma'} - C_{\gamma})RT \ln(C_{\gamma}/C_m)}{(1 - C_{\gamma})[1 + \partial \ln v / \partial \ln C_{\gamma}]}, \quad (15a)$$

A similar relation applies when the precipitate is depleted in solute (e.g., Cr, Co), i.e.,

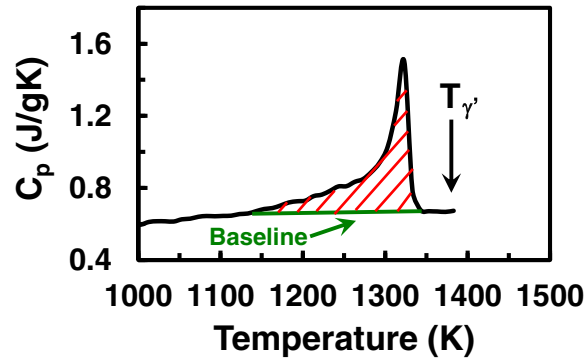


Fig. 2 Schematic illustration of the determination of the enthalpy associated with the decomposition of supersaturated γ to form $\gamma + \gamma'$ from on-cooling C_p measurements. The pertinent enthalpy is the cross-hatched region

$$\Delta G^*(\gamma \rightarrow \gamma') = - \frac{(C_Y - C_{Y'})RT \ln[(1 - C_Y)/(1 - C_m)]}{C_Y [1 + \partial \ln v / \partial \ln C_Y]} \quad (15b)$$

in which all of the terms are the same as defined above.

Values of ΔG^* for an undercooling of 30 K below the solvus temperature have been calculated for a supersaturated solid solution produced by supersolvus heat treatment of IN-100, René 88, and LSHR; i.e., C_m = solute content specified by the overall alloy composition. The calculations utilized Eq. (15) with the alloy/phase composition data in Tables 2 and 3 and estimates of the thermodynamic factor for each alloying element from the commercial code Pandat™ (CompuTherm LLC, Madison, WI) (Table 5). The *equilibrium* solute composition in the matrix (C_Y) was determined from the precipitate composition (Table 3) and the equilibrium volume fraction of γ' per the solvus approach curves (Fig. 1). The results of the calculations (Table 5) reveal that the largest values of ΔG^* correspond to those for chromium for all of the alloys. As suggested by the composition data in Tables 2 and 3, this behavior may be ascribed to a very high level of solute partitioning between the γ and γ' phases for this element.

The importance of chromium in controlling the driving force for nucleation is confirmed by a comparison of the predicted variation of ΔG^* with temperature from the methods based on either Eq. (13) or Eq. (15b) (for Cr solute) for a supersaturated solid solution produced by supertransus solution treatment. The comparison for IN-100 and LSHR shows excellent agreement (Fig. 4a). Although not formally appropriate, predictions from Eq. (15a), also plotted in Fig. 4a, show very good agreement with those from Eqs. (13) and (15b) as well. Furthermore, a comparison of the ΔG^* dependence on temperature for supersolvus-solution-treated alloys LSHR and ME3, derived using Eq. (15b), shows essentially identical results. This behavior is as expected based on the very similar alloy and γ' compositions, in particular with respect to the Cr content (Tables 2 and 3).

For the general case of an arbitrary solution temperature and non-zero values of the precipitate volume fraction, Eq. (15) is readily applied to determine ΔG^* and is thus typically used in fast-acting simulations.

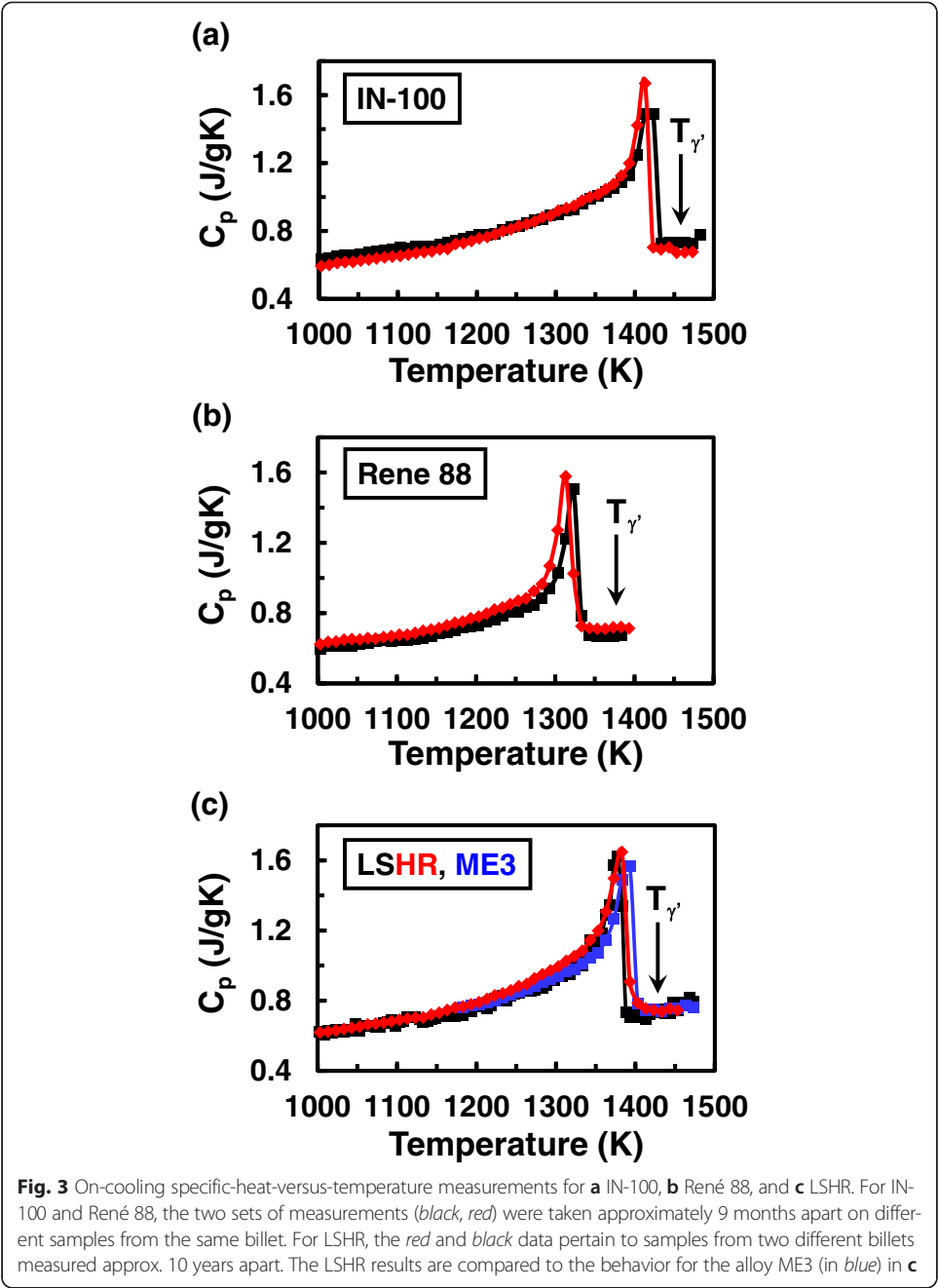


Table 4 Average enthalpy and entropy of a supersaturated solid solution determined from specific-heat data and gamma-prime solvus

Material	$T_{\gamma'}$ (K)	ΔH_{avg} (J/mol)	ΔS_{avg} (J/molK)
IN-100	1458	7045	4.83
René 88	1378	5807	4.21
LSHR	1430	7206	5.04

Table 5 ΔG^* (J/mol) for Ni-X pseudo-binary alloys for an undercooling of 30 K below T_γ

Element, X	IN-100	René 88	LSHR
Co	79.7	62.6	80.4
Cr	155.7	95.1	139.7
Al	16.9	38.4	21.4
Ti	26.5	66.8	32.1
Nb	–	22.6	7.0
Ta	–	–	6.6
Mo	12.0	3.2	7.3
W	–	1.1	0.1

Retarding force for nucleation

The principal retarding force for nucleation consists of the creation of γ - γ' interfaces with their associated interfacial energy σ . The determination of σ is usually based on the temperature at which “noticeable” nucleation occurs. This onset temperature per se can be deduced using a variety of techniques including in situ (synchrotron) x-ray diffraction (e.g., reference [50]), differential thermal analysis (DTA), and on-cooling specific-heat measurements. The application of the latter technique is shown in Fig. 3, which reveals the temperature/undercooling at which the specific heat begins to rise

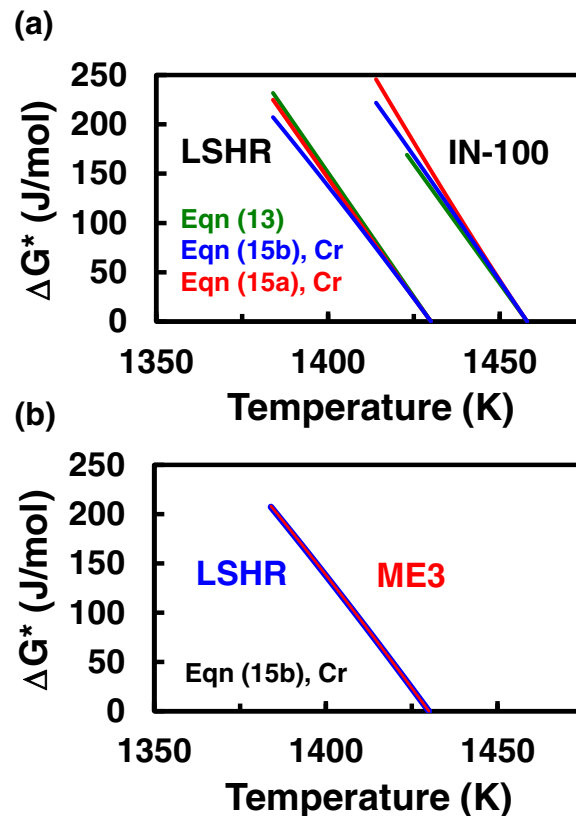


Fig. 4 Temperature dependence of ΔG^* for nucleation of γ' for supersolvus-solution-treated PM superalloys: **a** results for IN-100 and LSHR predicted using Eq. (13), (15b), or (15a), the latter two relations based on Cr partitioning and **b** comparison for LSHR and ME3 using Eq. (15b) based on Cr partitioning. The results in **b** indicate negligible difference for the two alloys

noticeably. These temperatures are found to be 1431, 1345, and 1403.5 K (corresponding to an undercooling of 27, 33, or 26.5 K), respectively, for IN-100, René 88, and LSHR/ME3. The onset temperature and critical undercooling for LSHR from the specific heat are essentially identical to the values determined previously for LSHR via the synchrotron method [50]. Furthermore, the previous in-situ results indicated that the onset temperature is identical for cooling rates of 11 and 139 K/min.

Following an approach similar to that suggested by Doherty [4], the surface energy is chosen to yield a value of J_o (per Eq. (1) with a diffusivity based on static-coarsening measurements, as discussed below) that is discernable by post heat-treatment metallography. For cooling rates of the order of 0.2–2 K/s, a critical J_o of the order of $1/\mu\text{m}^3\text{s}$ ($10^{18}/\text{m}^3\text{s}$) at the observed nucleation-onset temperature is reasonable. Examples of the determination of σ for IN-100, René 88, and LSHR by this means are summarized in Fig. 5. The surface energies so deduced are between 23 and 25 mJ/m^2 . Because of the exponential dependence of nucleation rate on σ^3 , the use of a different value for the

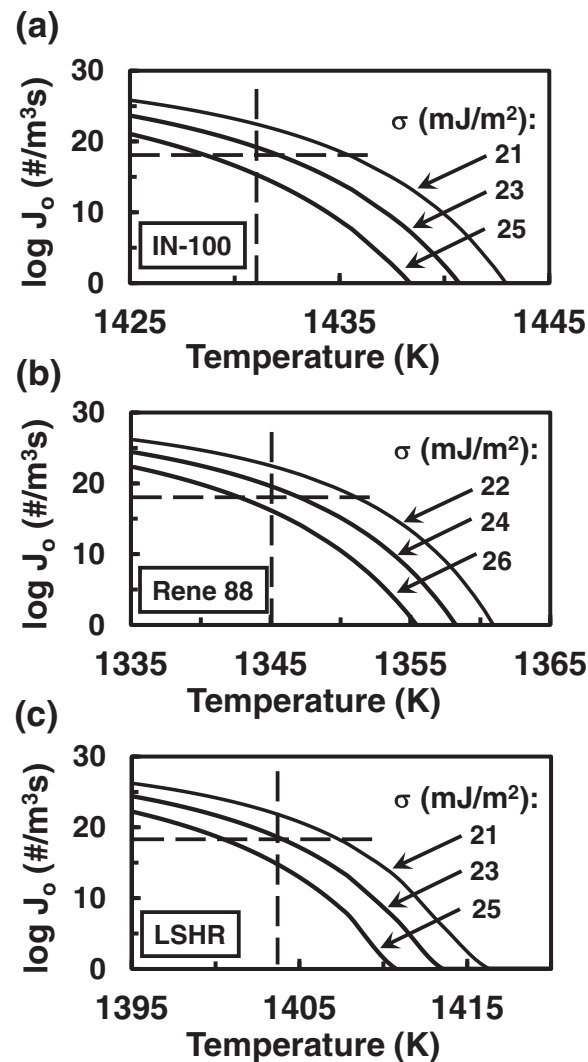


Fig. 5 Plots of the calculated steady-state nucleation rate (J_o) as a function of temperature used to estimate the γ - γ' interface energy (σ) for **a** IN-100, **b** René 88, and **c** LSHR

“critical” value of J would have led to only a small difference in the appropriate choice of σ . For example, if the critical J were chosen to be $0.1/\mu\text{m}^3\text{s}$, the surface energy for the three alloys would have been between 24 and 26 mJ/m^2 per the plots in Fig. 5.

The values of σ in the present work are almost identical to that deduced by Sudbrack et al. [28] for a ternary Ni-Al-Cr alloy with levels of aluminum and chromium similar to those in IN-100 and LSHR. By contrast, the present value is considerably lower than the value quoted by Olson et al. [35] for LSHR (i.e., 31.5 mJ/m^2). Perhaps, this difference can be rationalized in the context of values of ΔG^* which may have been too high in the earlier prior work. For example, an examination of Eqs. (1) and (2) indicates that the nucleation rate depends on an exponential term whose argument includes a factor of $\sigma^3/\Delta G^{*2}$. Thus, to obtain an identical/observable nucleation rate at a given temperature, a surface energy that was high by a factor of approximately 35 pct would have required that ΔG^* to be overestimated by 57 pct.

The misfit energy ΔG_p can also act as a retarding force through its effect in raising the overall system energy in much the same way that the interface energy does as precipitates are formed [55]. However, in most cases, ΔG_p is small relative to that of ΔG^* and can be neglected to a first order or included implicitly with ΔG^* . For example, Booth-Morrison et al. [56] determined that the misfit was ~ 0.2 pct for the ternary alloy Ni-5.2Al-14.2Cr (in atomic percent), i.e., a material whose levels of chromium and aluminum are similar to those in commercial PM superalloys. For this level of misfit, ΔG_p would be of the order of 1 J/mol. Such a value is two orders of magnitude smaller than the values of ΔG^* for temperatures at which secondary and tertiary γ' nucleation occurs.

Effective diffusivity, D_{eff}

The diffusivities of various solutes in nickel-base superalloys play a very important role in the precipitation process through their effect on particle growth per se as well as the rate of depletion of the matrix supersaturation controlling nucleation behavior. Because alloying elements such as Al and Ti partition in a sense opposite to that of Cr in the γ and γ' phases and the off-diagonal terms in the diffusivity matrix are non-zero, the development of a concentration gradient for one alloying element may noticeably retard the overall diffusive flux of another. For example, Al and Cr in a nickel solid solution have a positive interaction such that Al can diffuse down a chromium concentration gradient [57–59]. Thus, as a γ' precipitate grows, the rate of diffusion of a given element to or away from the particle may be mitigated somewhat by its tendency to diffuse down the concentration gradient of a different alloying element.

Because of the complexity of the diffusion problem, it is therefore often simpler to determine an *effective* diffusivity for the alloying element whose behavior appears to be rate limiting and to which simple (pseudo-binary) diffusion analyses can be applied. For PM nickel-base superalloys such as those of interest here, the work of Campbell et al. [60] and Semiatin et al. [50] suggests that Cr diffusion is rate limiting in γ - γ' superalloys.

In the present work, the specific method used to fit an effective diffusivity has been based on an analysis of the rate of coarsening of *intragranular*, secondary γ' during isothermal heat treatments. In particular, experimental observations (e.g., Fig. 6 for IN-100 and Fig. 7 for René 88) were interpreted using Eq. (10), assuming that the process

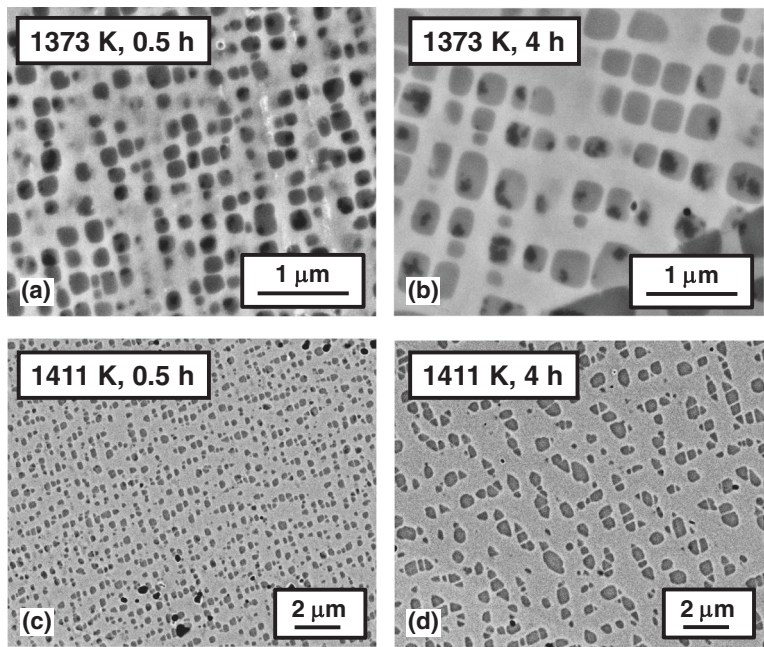


Fig. 6 Backscatter-electron images of IN-100 illustrating static coarsening of γ' at **a, b** 1373 K or **c, d** 1411 K

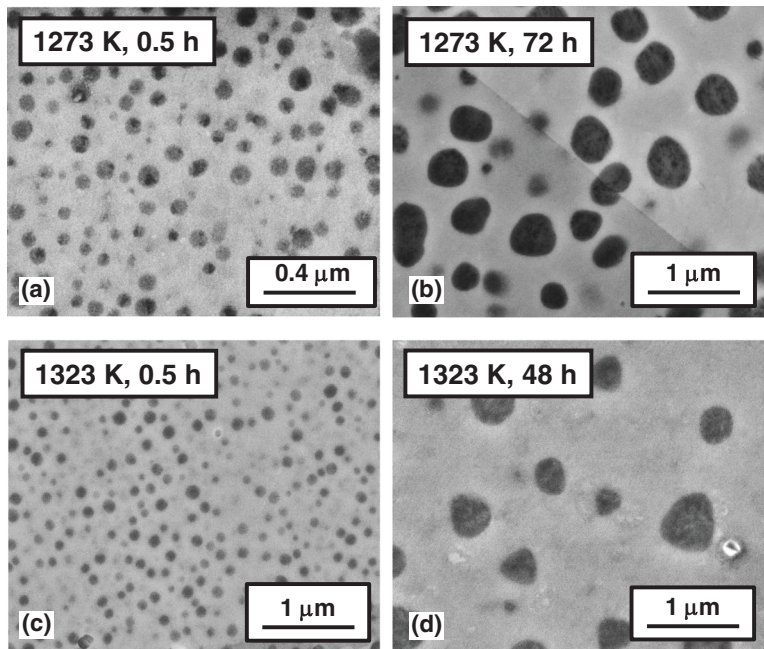


Fig. 7 Backscatter-electron images of René 88 illustrating static coarsening of γ' at **a, b** 1273 K or **c, d** 1323 K

has been controlled solely by Cr diffusion, or Eqs. (10) and (11) to account for the influence of all of the solutes. In both analyses, the effective diffusivity of Cr as a function of temperature was taken equal to the product of the impurity diffusivity, D_B , of Cr in binary Ni-Cr alloys [61] (principally to obtain the activation energy/temperature dependence) and a fitting factor, A , to account for the finite concentration of Cr in the γ matrix and its interaction with other solutes, i.e.,

$$D_{\text{eff}}(\text{m}^2/\text{s}) = A \times D_B = A \times 0.00036 \exp(-34,278/T(K)). \quad (16)$$

The diffusivities of the various other solutes in each alloy were taken to be in the ratios (relative to that of Cr) suggested by the work of Campbell et al. [60]. In addition, $w(f)$ was taken from the research of Voorhees and Glicksman [20], and σ was assumed to be 23 mJ/m² per the results in Fig. 5. The molar volume V_M was calculated to be 7.20×10^{-6} , 7.11×10^{-6} , and 7.22×10^{-6} m³/mol for IN-100, René 88, and LSHR, respectively. The thermodynamic factors needed to apply Eq. (10) were derived from Pandat™ calculations (Table 6).

Predicted coarsening rate constants for IN-100, René 88, and LSHR at various temperatures and two different values of A ($= D_{\text{eff}}/D_B$), 0.33 and 0.8, are summarized in Table 7. For a given alloy, temperature, and value of A , the calculations reveal that the predicted coarsening rate based on Cr diffusion alone (Eq. (10)) is approximately 1.5 to 3 times that determined when accounting for the interaction among the solutes using Eqs. (10) and (11). Table 7 also summarizes *measured* rate constants based on SEM observations on 2D sections (e.g., those in Figs. 6 and 7). Two different methods were used to reduce such measurements. In one case (i.e., annotated as “SC”, or “stereological correction”), it was assumed that the true 3D diameters were 15 pct larger than those measured on the 2D sections. In the other case, no such SC was applied. This latter approach has been justified based on recent geometric analysis by Payton et al. [62] for distributions of spherical particles that have finite breadth, as is pertinent for precipitates undergoing coarsening. Specifically, it was deduced that the SC used in the first case *overestimates* the actual average 3D size and that no correction factor is actually needed for distribution shapes such as those developed during static coarsening.

Taken as a whole, the comparison of predicted and measured values of the coarsening rate constant (Table 7) shows two important trends. First, $A = 0.8$ is more appropriate for IN-100 and René 88, but a lower value, $A = 0.33$, gives a better fit for LSHR.

Table 6 Thermodynamic factors (thermodynamic factor = $1 + \partial \ln v / \partial \ln C_\gamma$) for solutes in gamma solid solution

Material	Temp (K)	Co	Cr	Al	Ti	Nb	Ta	Mo	W
IN-100	1430	1.02	1.2	2.9	2.3	–	–	0.97	–
René 88	1083	1.05	1.72	1.65	1.52	1.05	–	1.0	1.10
René 88	1227	1.06	1.52	1.71	1.65	1.06	–	1.0	1.09
René 88	1323	1.03	1.41	1.79	1.95	1.08	–	0.99	1.09
René 88	1343	1.03	1.33	1.93	2.18	1.12	–	0.98	1.10
LSHR	1116	1.06	1.54	2.11	1.28	1.09	1.02	0.97	1.09
LSHR	1200	1.05	1.46	2.19	1.42	1.11	1.04	0.97	1.09
LSHR	1366	0.93	1.26	2.5	1.97	1.25	1.2	0.96	1.14
LSHR	1403	0.90	1.18	2.15	1.8	1.20	1.18	0.97	1.10

Table 7 Application of coarsening data to calibrate the effective diffusivity

Mat'l	Temp (K)	v/o γ'	w(f)	D_{eff}/D_B	Pred K, Cr ($\mu\text{m}^3/\text{s}$)	Pred K, All ($\mu\text{m}^3/\text{s}$)	Meas w SC ($\mu\text{m}^3/\text{s}$)	Meas w/o SC ($\mu\text{m}^3/\text{s}$)
IN-100	1373	0.38	2.75	0.33	4.88×10^{-7}	1.55×10^{-7}	4.55×10^{-7}	2.99×10^{-7}
IN-100	1373	0.38	2.75	0.80	11.8×10^{-7}	3.75×10^{-7}	4.55×10^{-7}	2.99×10^{-7}
IN-100	1411	0.20	2.0	0.33	6.48×10^{-7}	4.02×10^{-7}	1.61×10^{-6}	1.06×10^{-6}
IN-100	1411	0.20	2.0	0.80	1.57×10^{-6}	0.98×10^{-6}	1.61×10^{-6}	1.06×10^{-6}
René 88	1273	0.20	2.0	0.33	2.54×10^{-8}	1.04×10^{-8}	2.75×10^{-8}	1.81×10^{-8}
René 88	1273	0.20	2.0	0.80	6.16×10^{-8}	2.53×10^{-8}	2.75×10^{-8}	1.81×10^{-8}
René 88	1323	0.13	1.7	0.33	6.78×10^{-8}	3.11×10^{-8}	9.38×10^{-8}	6.17×10^{-8}
René 88	1323	0.13	1.7	0.80	16.4×10^{-8}	7.54×10^{-8}	9.38×10^{-8}	6.17×10^{-8}
LSHR	1366	0.19	1.9	0.33	1.81×10^{-7}	1.14×10^{-7}	1.83×10^{-7}	0.99×10^{-7}

SC stereological correction

Second, the predicted rate constants based on the interaction of all of the solutes are closer to the measurements than those based on diffusion of Cr alone. In fact, the agreement is remarkably good for the various alloys and different temperatures when the calculation incorporates the effect of all of the solutes and uses $A = 0.8$ for the first- and second-generation alloys and $A = 0.33$ for the third-generation alloy. The generality of this conclusion has been confirmed in a companion report [63] containing additional coarsening data and analysis for three third-generation PM alloys (LSHR, ME3, Alloy 10). In this other effort, the value of A was deduced to be 0.33 for all of three alloys, and, not surprisingly, the measured coarsening rates for LSHR and ME3 (having similar alloy and γ' compositions) were identical. An inspection of the compositions of IN-100/René 88 and the various third-generation alloys (e.g., Tables 1 and 2) suggests that the tantalum addition is the differentiating element which plays a critical role in reducing the rate of diffusion which controls coarsening and likely the kinetics of precipitate nucleation and growth. In the latter regard, it was shown in reference [50] that $A = 0.33$ provides excellent predictions of the size and number density of secondary γ' developed during continuous cooling of LSHR following supersolvus solution treatment.

Despite the generality of the present findings regarding the effective diffusivity, a word of caution is in order. That is to say, the companion effort [58] has shown that calculations assuming bulk-diffusion control, as embodied in Eq. (10), *overestimate* the measured coarsening rate by approximately a factor of five for temperatures *below* approximately 1050 K or those typical of service conditions for PM nickel-base superalloys. At such temperatures, the coarsening mechanism changes from bulk-diffusion control to more-sluggish trans-interface-diffusion control [64].

Summary

Engineering methods for determining the input parameters for simulating nucleation, growth, and coarsening of γ' in PM nickel-base superalloys have been formulated and validated for IN-100, René 88, and LSHR. The simplicity of the techniques makes them excellent candidates for inclusion in ICME “best-practice” documents. The approaches can be summarized as follows:

1. *Solvus approach curve and phase compositions*: The solvus approach curve can be fit by an analytical expression with three parameters—the solvus temperature per

se, the maximum amount of γ' based on the atomic fractions of γ' -forming elements, and a parameter Q used to fit the expression to a relatively small number of experimental measurements. The value of Q lies in the range between 60 and 75 kJ/mol. Phase compositions can be determined using conventional phase-extraction methods and subsequent chemical analysis of each phase.

2. The *chemical free energy* associated with the $\gamma \rightarrow \gamma + \gamma'$ transformation is readily determined from specific-heat/solvus temperature measurements or classical expressions from solution thermodynamics for the Ni-Cr pseudo-binary. The latter method is suitable for arbitrary matrix compositions, solution temperatures (super- or subsolvus), and complex cooling paths.
3. The γ - γ' *interface energy* σ is easily determined from the temperature at the onset of nucleation, simple parametric calculations of the nucleation rate (in which σ is varied), and the specification of an “operational” critical nucleation rate (e.g., $1/\mu\text{m}^3\text{s}$). The onset temperature is readily determined from careful on-cooling measurements of the specific heat. The value of σ for PM superalloys analyzed in this work has been found to lie in a narrow window between 23 and 25 mJ/m².
4. An *effective diffusivity* based on the impurity diffusion for chromium in nickel and a scaling factor “ A ” independent of temperature can be readily determined from static-coarsening measurements at several temperatures. The scaling factor has been found to be approximately 0.8 for IN-100 and René 88 and 0.33 for third-generation PM superalloys which contain tantalum (i.e., LSHR, ME3, and Alloy 10).

List of symbols

A ,	scaling factor for effective diffusivity
a_0 ,	lattice parameter
C ,	concentration
$C_\gamma, C_{\gamma'}$,	equilibrium solute concentration in the γ matrix, γ' precipitate
C_p ,	specific heat
C^* ,	total atomic fraction of γ' -forming elements
D ,	diffusivity
D_{eff} ,	effective diffusivity
D_B ,	impurity diffusivity of solute in nickel
f ,	volume fraction of precipitate
J ,	nucleation rate
J_o ,	steady-state nucleation rate
K_{MLSW} ,	volume-fraction modified LSW coarsening-rate constant
k_B ,	Boltzmann's constant
Q ,	fitting parameter for solvus approach curve
r ,	precipitate radius
r^* ,	critical radius of precipitate
R ,	gas constant
T ,	absolute temperature
$T_{\gamma'}, \gamma'$	solvus temperature
t ,	time
V_M ,	molar volume
ν ,	activity coefficient

w , volume-fraction function in coarsening-rate equation

ΔG^* , chemical free energy of transformation

ΔG_p , misfit energy

ΔH , enthalpy of formation of γ'

ΔS , entropy of formation of γ'

Ω , supersaturation

σ , matrix-precipitate interface energy.

τ , incubation time

Endnotes

¹Heterogeneous precipitation of γ' at γ grain boundaries, which can play a role in creep and dwell fatigue during service, is not treated herein.

Competing interests

The authors declare that they have no competing interests.

Authors' contributions

SLS led the R&D and developed the theoretical analysis methods. FZ performed the thermodynamic modeling including the determination of thermodynamic factors. RL and LAC developed the test techniques and data analysis methods for determining the on-cooling specific heat and pertinent transformation temperatures. DUF guided the definition of the program scope and its ICME application and mentored the other team members. All authors read and approved the manuscript.

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Author details

¹Air Force Research Laboratory, Materials and Manufacturing Directorate, AFRL/RXCM, Wright-Patterson Air Force Base, OH 45433-7817, USA. ²CompuTherm, LLC, Madison, WI 53719, USA. ³Thermophysical Properties Research Laboratory, Inc., West Lafayette, IN 47906, USA. ⁴National Physical Laboratory, Teddington, Middlesex TW11 0LW, UK. ⁵Pratt & Whitney, 400 Main Street, East Hartford, CT 06118, USA.

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